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WORKSHEET # 6 MIXED MOLE

PROBLEMS (GRAMS, MOLECULES, AND

LITERS) You now know three things a

mole can be: a molar mass, 6.02×10^{23}

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molecules and, for a gas, 22.4 liters at STP.

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WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find

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the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

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the numbers in the Mole Concept

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While a dozen is only 12 particles a mole is a much larger number— 6.02×10^{23} particles. Elements generally exist as the particles we call atoms. A mole of carbon contains 6.02×10^{23} atoms of carbon. A mole of helium contains $6.02 \times$

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1023atoms of helium. A mole of sodium contains 6.02×10^{23} atoms of sodium.

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TO MOLES. Molar Mass (# grams) 6.02×10^{23} ATOMS (for elements) or $23 \times 6.02 \times 10$ MOLECULES (for compounds) Atomic Weight (for elements) or Gram Formula/Molecular Weight (for compounds) MOLES 22.4 LITERS OF A GAS AT STP (standard temperature and pressure) FORMULA CALCULATION MASS

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CHEMISTRY WORKSHEET # 2: THE MOLE
AS A UNIT OF ...

chemistry worksheet # 2: the mole as a unit of mass

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WORKSHEET # 2: THE MOLE AS A UNIT
OF MASS Now that you know how to find
the mass of one mole of a substance
(molar mass) you can easily find the
mass of several moles or the mass of a
fraction of a mole using the factor-label
technique. 1 mol = a molar mass of an
atom/molecule (g/mol) EXAMPLE: What is

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the mass of 5.00 moles of water H ...

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A) 8.5 moles of sodium acetate dissolved in 1 L of water B) 5.5 moles of sodium acetate dissolved in 500 mL of water C) 1.8 moles of sodium acetate dissolved in 300 mL of water D) 1.2 moles of sodium acetate dissolved in 200 mL of water E)

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none of the above

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